

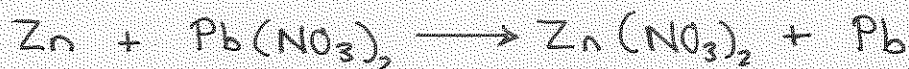
Practice Sheets

WORD EQUATIONS

Name _____

Write the word equations below as chemical equations and balance.

1. zinc + lead (II) nitrate yield zinc nitrate + lead



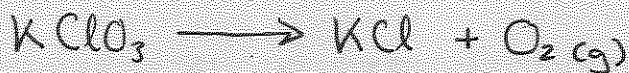
2. aluminum bromide + chlorine yield aluminum chloride + bromine



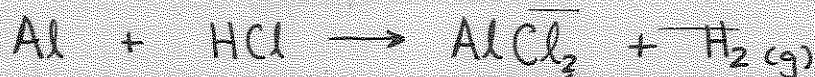
3. sodium phosphate + calcium chloride yield calcium phosphate + sodium chloride



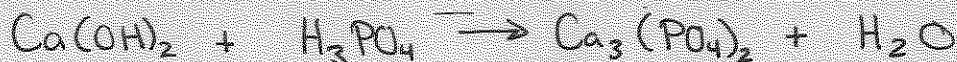
4. potassium chlorate when heated yields potassium chloride + oxygen gas



5. aluminum + hydrochloric acid yield aluminum chloride + hydrogen gas



6. calcium hydroxide + phosphoric acid yield calcium phosphate + water



7. copper + sulfuric acid yield copper (II) sulfate + water + sulfur dioxide



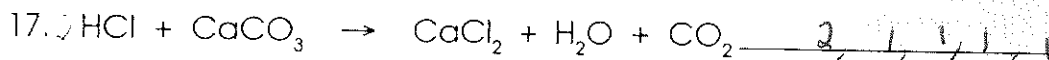
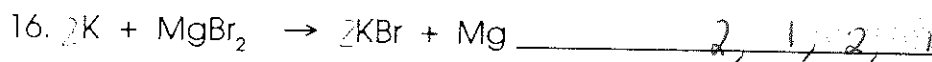
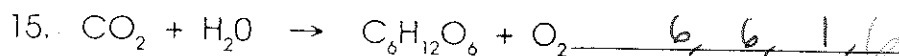
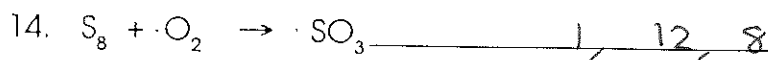
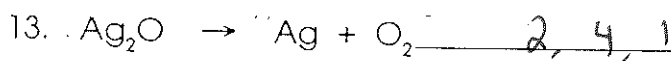
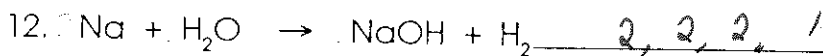
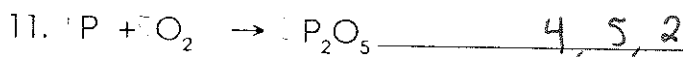
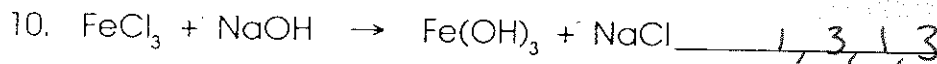
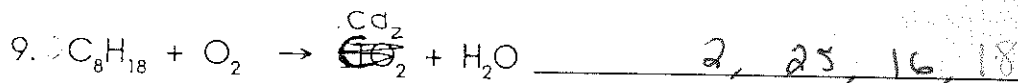
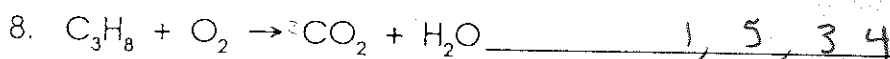
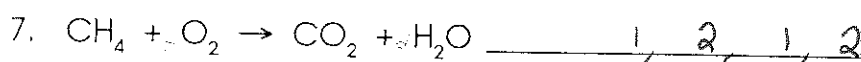
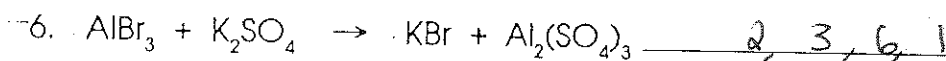
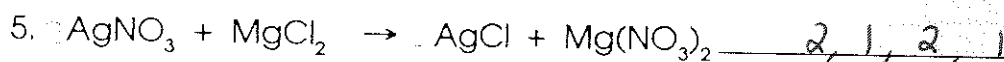
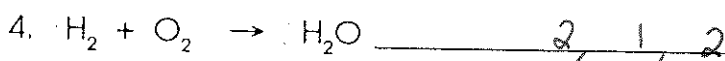
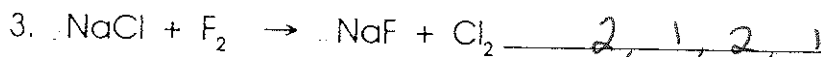
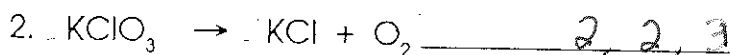
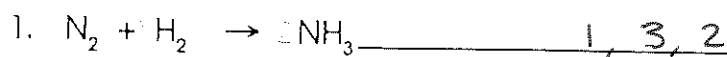
8. hydrogen + nitrogen monoxide yield water + nitrogen



BALANCING CHEMICAL EQUATIONS

Name _____

Rewrite and balance the equations below.



CLASSIFICATION OF CHEMICAL REACTIONS



Name _____

Classify the reactions below as synthesis, decomposition, single replacement (cationic or anionic) or double replacement.



synthesis



decomposition



single replacement



synthesis



decomposition



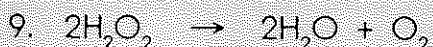
single replacement



synthesis



double replacement



decomposition



double replacement

PREDICTING PRODUCTS OF CHEMICAL REACTIONS

Name _____

Predict the products of the reactions below. Then, write the balanced equation and classify the reaction.

1. magnesium bromide + chlorine



2. aluminum + iron (III) oxide



3. silver nitrate + zinc chloride



4. hydrogen peroxide (catalyzed by manganese dioxide)



5. zinc + hydrochloric acid



6. sulfuric acid + sodium hydroxide



7. sodium + hydrogen



8. acetic acid + copper



Write whether the Reaction is Exothermic or Endothermic

