

Redox Practice (Balancing Redox)

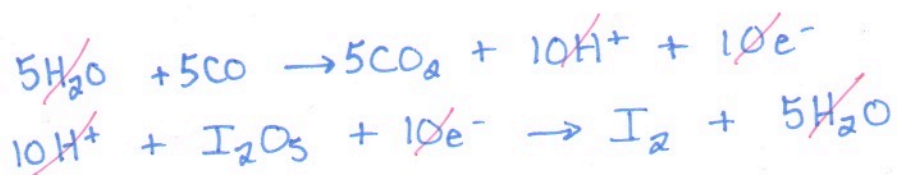
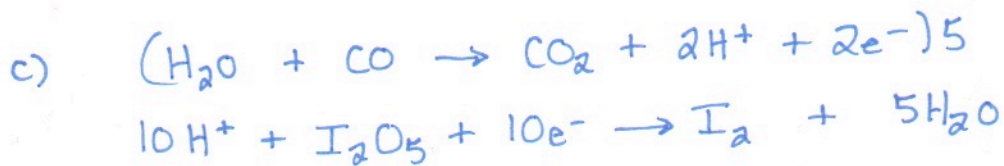
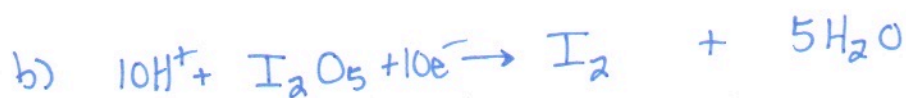
1. Consider the following redox equation: $\text{CO} + \text{I}_2\text{O}_5 \rightarrow \text{I}_2 + \text{CO}_2$

Write a balanced equation for each of the following: (4 marks)

- oxidation half-reaction.
- reduction half-reaction.
- redox equation.

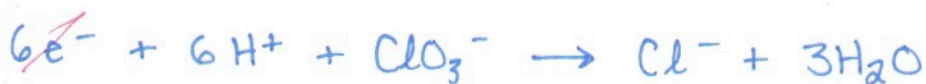
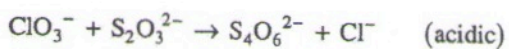
a)

can't balance the half reactions without H^+ & O ,
therefore treat it like it's acidic



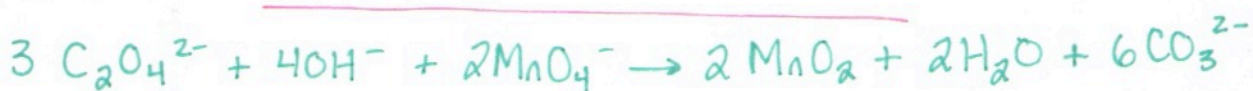
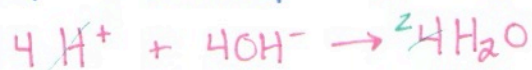
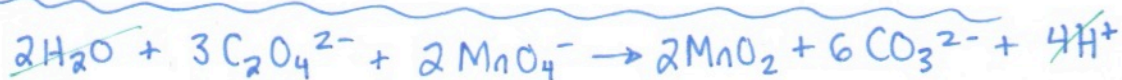
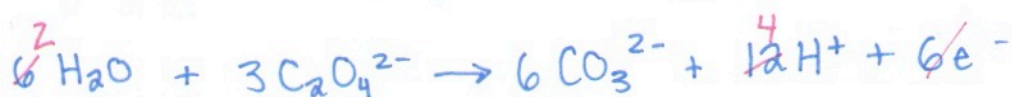
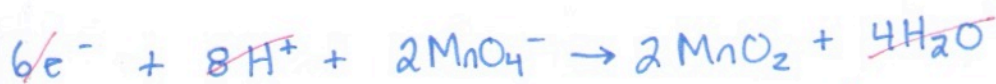
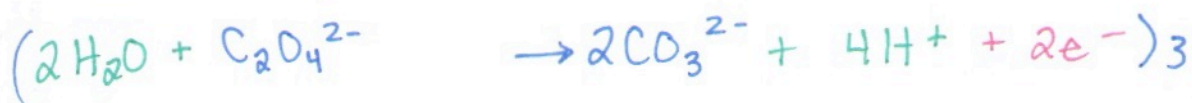
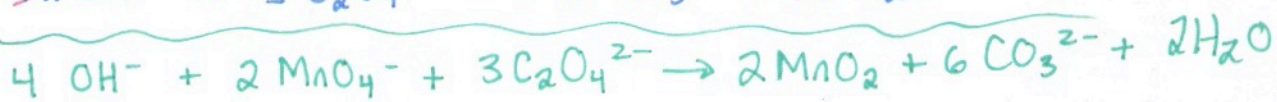
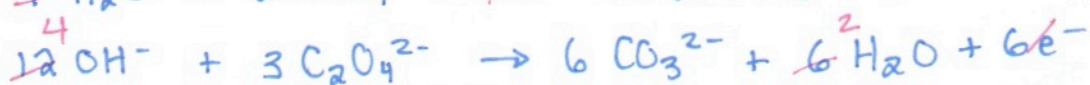
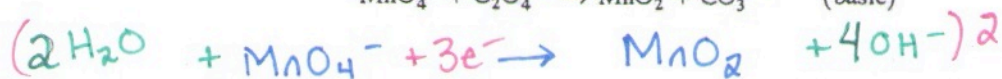
2. Balance the following redox equation:

(4 marks)



3. Balance the following redox reaction in basic solution.

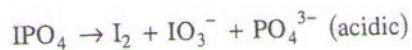
(5 marks)



A hard one:

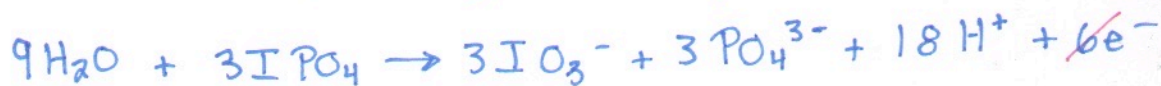
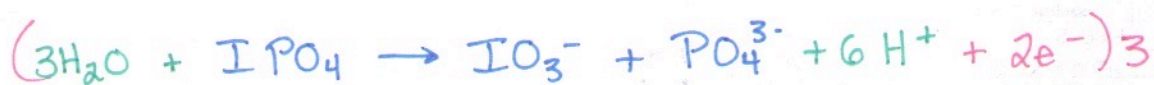
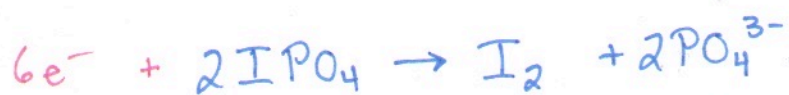
4.

In an unusual compound, IPO_4 , iodine exists as Iodine(III). The compound decomposes as follows:



Balance this redox equation in acidic solution.

(4 marks)



2 I

